

Supplementary texts: *These textbooks are available in A-204*

Chemistry 0110 Review by Fraser

Chemistry: A Study of Matter by Dorin

Introductory Chemistry: Zumdahl

Basic Chemistry by Seese and Daub 7th Edition

Chemistry: A Basic Introduction by Miller 4th Edition

CALENDAR

DESCRIPTION: The major concepts covered include: inorganic nomenclature; modern atomic structure, orbitals; ionic and covalent bonding, hydrogen bonding, metallic bonding, Van der Waal forces, ionization, electronegativity, VSEPR; solutions, stoichiometry, empirical formulas, percent composition, pH, molarity, equilibrium, Arrhenius acids and bases.

Credit/Contact Hours: This is a 5 credit course and meets 6 hours per week (4 hour lecture and 2 hours lab).

DELIVERY MODE(S): Lecture will be the main method of delivery. There is also a large laboratory component in this course.

ASSUMED BACKGROUND KNOWLEDGE: (These topics will **NOT** be reviewed in class)

Students should already:

1. be able to perform linear, quadratic, cubic, liquid-dry metric conversions.
2. know elementary atomic structure – what protons, electrons, and neutrons are and where they are found in an atom.
3. be able to define atomic number, atomic mass number and how to use them to calculate the number of protons, neutrons, and electrons in an atom.
4. be able to classify of matter.
5. be able to distinguish between chemical and physical properties and changes.
6. be able to draw atomic structure diagram for the first 20 elements.
7. be able to define valence electrons and draw electron dot diagrams.

If you are unfamiliar with these topics see me. The Chemistry 0110 Review will explain these in detail. It is available in the bookstore. There is an answer key in A-205 and on reserve in the library.

In addition to the above material: These will be briefly reviewed in class.

1. know nomenclature (naming compounds and writing formulae). ***Nomenclature is one of the most important topics that you will learn at the secondary level. It will NOT be reviewed at the post secondary level. If you are having trouble with this topic, get help IMMEDIATELY!!! See me! **In this course you will be expected to do nomenclature without a periodic table*****
2. be able to balancing equations by inspection.

OBJECTIVES:

Students should

1. be able to do nomenclature **without** the use of a periodic table.
2. be able to define temperature, pressure, vapour pressure, and boiling point.
3. be able to define solubility, and state factors affecting solubility and the rate of solution.
4. be able define unsaturated, saturated, supersaturated solutions and interpret solubility curves.
5. be able to solve a variety of stoichiometry problems.
6. be able to calculate the % composition of each element in a compound.
7. be able to define and find empirical formulae given % composition by mass of each element.
8. be able to calculate percent error and percent yield.
9. understand the kinetic molecular theory of gases.
10. know and define gas laws (Boyle's Law, Charles' Law, combined gas laws, the ideal gas law.) Volume of gases under STP conditions and SATP if time permits.
11. be able to solve stoichiometry problems from chemical equations including determining the limiting reagent.
12. understand equilibrium, write equilibrium equations for given reactions and understand the effect of changing concentration.

13. be able to write solubility product expression for given compounds.
14. understand and define K_w and pH and to able to perform related calculations.

*****Midterm*****

15. be able to write the electronic and orbital box diagrams for any element.
16. be able to compare the reactivity and radius based on atomic and electronic structures, ionization energy and electronegativities.
17. be able recognizes trends in the periodic table.
18. be able to draw the structural diagrams for various molecules etc. using VSEPR.
19. be able to distinguish between polar, nonpolar covalent and ionic bonds by calculating the difference in electronegativities and observing the shapes of molecules.
20. be able to recognize and define a coordinate covalent bond.
21. be able to recognize and define polar, and nonpolar molecules from the diagrams.
22. be able to recognize and define hydrogen bonding, and Van der Waal's force.

TRANSFERABILITY: This course is equivalent to Alberta grade 11 chemistry and is transferable to other post secondary institutions.

GRADING CRITERIA: Regular attendance is expected of all students, and is crucial to passing this course. Students who miss classes will soon find themselves falling behind and failing. Lateness will **not** be tolerated as it interrupts the instructor and fellow classmates. As per Department Policy, if you miss more than 10 per semester of classes in any course, you may be debarred from the final exam for that course.

A certificate (a doctor's or a note from the funeral home) will be required to make up the midterm or final exam. **You will receive a grade of F if you miss the final.** Call if you are going to miss a test. There may be a deduction of 10% for test rewrites.

*****Very important:**

Laboratory attendance to each specific experiment is compulsory; a passing grade in the laboratory component is required to pass the course. There are **NO** ‘make up’ labs in this course. Being absent from an experiment will result in a grade of **ZERO** for that experiment.

Lab reports must be submitted on the required date and at the **required time**.

Assignments may not be accepted after the assignment has been returned to the class. I am usually a speedy marker and usually return papers the next day.

Penalties for late **assignments** are as follows: (Assuming that I have not returned the marked assignments)

1 day late – 20%, 2 days late – 50%, 3 days late – 100%

Penalties for late **lab reports** are as follows:

5 minutes after due time - 10 %, 24 hours after due time - 100%

Marking Scheme:

Lab Reports:	15%
Assignments:	15%
Tests:	15%
Midterm:	15%
Final Exam:	<u>40%</u>
Total	100%

GRANDE PRAIRIE REGIONAL COLLEGE			
GRADING CONVERSION CHART			
Alpha Grade	4-point Equivalent	Percentage Guidelines	Designation
A⁺	4	90 – 100	EXCELLENT
A	4	85 – 89	
A⁻	3.7	80 – 84	FIRST CLASS STANDING
B⁺	3.3	77 – 79	
B	3	73 – 76	GOOD
B⁻	2.7	70 – 72	
C⁺	2.3	67 – 69	SATISFACTORY
C	2	63 – 66	
C⁻	1.7	60 – 62	
D⁺	1.3	55 – 59	MINIMAL PASS
D	1	50 – 54	
F	0	0 – 49	FAIL
WF	0.0	0	FAIL, withdrawal after the deadline

EXAMINATIONS: See asterisks on the time line table.

- * Test 1
- ** Test 2
- *** Test 3 If students want a test on equation stoichiometry.
- *** Midterm ***
- **** Test 4

Final exams are set by the Registrar's Office

STUDENT RESPONSIBILITIES:

Students will:

- review material that is prerequisite to this course quickly so it does not slow you down. (See Assumed Background Knowledge on pages 2 & 3. Especially the section on nomenclature.)
- be at class regularly and on time. (If you miss more than 10 per semester of classes in any course, you may be debarred from the final exam for that course.)
- complete all pre class and pre-lab assignments before arriving in class.
- keep up with course material.
- if experiencing difficulties with course get help immediately.
- catch up on missed material before the next class.
- provide documentation for missed midterms or finals.
- be aware of penalty for failing the lab component and not writing the final.

STATEMENT ON PLAGIARISM AND CHEATING:

Please refer to pages 50 – 51 of the College calendar regarding plagiarism, cheating and the resultant penalties. These are serious issues and will be dealt with severely. The College calendar is available on line.

COURSE SCHEDULE/TENTATIVE TIME LINE:

	Time	Chemistry 0110 Review (Fraser)	Chemistry: A Study of Matter (Dorin) 4 Edition	General Chemistry Ebbing 8th Edition	Introductory Chemistry (Zumdahl) 5 th Edition	Basic Chemistry (Seese and Daub) 7 th Edition	Chemistry: A Basic Introduction Miller 4 th Edition
SI units	On your own time	19 – 30	16 – 22	20 – 22	18 – 22	16 – 18	19 – 23
Scientific Notation	On your own time	31 – 37	50 – 53		15 – 18	27 – 33	236 – 38
Understand and use Significant Figures	1 day	37 – 51	26 – 37 54 – 58	16 – 20	22 – 28	21 – 27	33 – 26 38 – 42
Chemical Definitions and Classification of matter	On your own time	14 – 18	63 – 77	9 – 13	56 – 64	52 – 61	79
WHMIS	On your own time	5 – 10					
Review Atomic Structure	On your own time	56 – 69	109 – 123 top of 764	42 – 53	214 – 216 91 – 97	80 – 89	60 – 72 90 – 102
Review Valence Electrons & Electron Dot Diagrams	On your own time	70 – 71				89 – 90	

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Review Nomenclature	2 days	79 – 150	151 – 168	55 – 71	122 – 142	165 – 180	81 – 82 188 – 183 197 – 206
*Review Balancing Equations by Inspection	On your own time	152 – 167	206 – 223	73 – 76 136 – 144 150 – 152	154 – 165	216 – 235	233 – 240
Solubility			437 – 438				
Solutions: Saturated. Unsaturated, Supersaturated Solution.	1 day		433 – 446	127 – 128		363 – 370 370 – 371	350 – 357
Factor Affecting Rates of Solution			442 – 439	479 – 496		363 – 370	357 – 363
Solutions: Concentrated and Dilute Solutions			446 – 447	157 – 159		379 – 381	

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Stoichiometry	<i>MOST IMPORTANT PART OF THIS COURSE!!!</i>						
Define & Calculate Molecular Mass, Gram Molecular Mass, Mole, Avogadro's Number Mole – Mass Relationship Mole – Molecule Relationship Mole-Volume of gas at STP Relationship	4 days		175 – 195 175 – 189 175 – 189	87 – 93 103 – 107	211 – 213 216 – 226 246 – 258	188 – 197	210 – 219
% Composition	1 day		190 – 191	93 – 94	226 – 228	199 – 205	219 – 223
**Empirical Formula	1 day		192 – 195	97 – 102	228 – 237		223 – 229
Kinetic Molecular Theory of Gases	1 day		271 – 283	201 – 204		277 – 278	280 – 281
Gas Laws: Boyle's Law, Charles' Law and Ideal Gas Law	2 days		297 – 313 326 – 328	176 – 190		279 – 288 285 – 297	285 – 302

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Volume of Gases under STP Conditions & (SATP Conditions If time permits)	1 day		297 – 313 326 – 328	176 – 190			303 – 308
Concentration (Molarity) (Molality)	1 day		449 – 451 452 – 453	155 – 159 161 – 163 491 – 492		377 – 379 384 – 386	353 – 368 372 – 373
***Equation Stoichiometry & Limiting Reagents	4 days		233 – 252	104 – 107	259 – 265	249 – 256 259 – 265	240 – 253
% Yield	if time			107 – 112	265 – 266	257 – 259	253 – 253
Equilibrium Constant	1 day		513 – 517 518 – 523	621 – 623		477 – 479	427 – 433
Le Chatelier Principle			524 – 527	639 – 642		480	433 – 440
Solubility Product (K_{sp})			533 – 537	735 – 744		490 – 493	440 – 444

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Arrhenius Acids, Bases, pH, & pOH (Strong acids and bases only)	1 day		550 – 552 558 – 559 574 – 580	660 – 662 673 – 681	185 – 188	406 – 407 418 – 422	389 – 390 397 – 409
Indicators			581 – 583	680 – 681			
Neutralization Reactions	1 day		584 – 586	136 – 139 140 – 142 161 – 164		414 – 417	
Atomic Structure Orbitals (s, p, d, f) Pauli Exclusion Principle & Hund's Rule	3 days		335 – 353	282 – 287 294 – 310		91 – 98	102 – 116
Ionization Energy & Electronegativity	2 days		369 – 373	316 – 318		119 130 – 132	136 – 139 140 – 142
****Trends in the Periodic Table	3 days		359 – 369 380 – 382	315 – 316 320 – 322			120 – 148

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Ionic Bonding	0.5 days		391 – 394	329 – 330		124 – 127	156 – 160
Covalent Bonding	0.5 days		396 – 402	341 – 347		128 – 135	161 – 164, 174
Hydrogen Bonding	0.5 days		416 – 419	440 – 442			
VSEPR Polar Molecules	5 days		402 – 409 412 – 416	347 – 350 373 – 380		135 – 147	164 – 180
Hydrogen Bonding	0.5 days		416 – 419	440 – 442			
Van der Waal Force	0.5 days		421 – 422	438 – 440			
Metal-Metal Bonding (If time permits)	0.5 days		419 – 420	bottom 534			155 – 156