



**Grande Prairie Regional College**  
**School of Business**  
**Department: Academic Upgrading**

**COURSE OUTLINE – FALL 2007**  
**PC0110 3(3-0-.5) HS Physics Grade 10 Equivalent**

<b>Instructor</b>	Medha Karnik	<b>Phone</b>	539-2952
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<b>Office</b>	Tuesday, Wednesday		
<b>Hours</b>	11:30 - 12:30 or by appointment		

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**PREREQUISITE:**

SC 0100; and MA 0110 (Acceptable as corequisite)

**TEXT BOOK:**

Locally developed modules by Dr Sekhon; are available at the Book Store. You may also use any other suitable text book to help you.

**CREDIT/CONTACT HOURS:**

PC 0110 is a 3.5 hr/week course with a small lab component.

**COURSE DESCRIPTION:**

The major concepts to be covered include kinematics; Newton's laws of motion; force; work, energy and power; and heat.

**ATTENDANCE AND LATENESS:**

Regular attendance is expected of all students which is crucial to passing the course. Students missing classes will soon find themselves falling behind and thus failing. Students with more than 15% absences (7 class days) may be barred from writing the final exam. Lateness is **highly discouraged**.

**ASSIGNMENTS, TESTS AND EXAMS:**

All tests and exams **MUST** be written on scheduled times. A missed test or exam will result in a score of zero unless **PRIOR** arrangements have been made with the Instructor for valid reasons to write the test/exam at some other time. All assignments **MUST** be handed by the deadline.

**LABS:**

There will only be a few labs in the course, and attendance in them is compulsory. A missed lab will result in a mark of zero. Makeup labs **CANNOT** be guaranteed, and may be permitted only under special circumstances. All labs reports **MUST** be handed in before the deadline. Late reports will result in severe penalties. Labs reports will **NOT** be marked if handed in late by more than two days unless pre-approval of the Instructor has been secured.

**EVALUATION:**

The course has been divided into four main units - Kinematics; Force; Work, energy, and power; and Heat (Thermal Energy). There will be an assignment and a unit test on each Unit, a Midterm exam, and the Final exam. The final grade will be based on different components as follows:

4 ASSIGNMENTS	=	20%
UNIT TEST	=	16%
MIDTERM EXAM	=	20%
FINAL EXAM	=	34%
LAB	=	10%

Grades will be assigned on the following Grading System.

**Academic Upgrading Department  
Grading Conversion Chart**

Alpha Grade	4-point Equivalent	Percentage Guidelines	Designation
A <sup>+</sup>	4	90 – 100	EXCELLENT
A	4	85 – 89	
A <sup>-</sup>	3.7	80 – 84	FIRST CLASS STANDING
B <sup>+</sup>	3.3	76 – 79	
B	3	73 – 75	GOOD
B <sup>-</sup>	2.7	70 – 72	
C <sup>+</sup>	2.3	67 – 69	SATISFACTORY
C	2	64 – 66	
C <sup>-</sup>	1.7	60 – 63	
D <sup>+</sup>	1.3	55 – 59	MINIMAL PASS
D	1	50 – 54	
F	0	0 – 49	FAIL



### **UNIT 3: WORK, ENERGY AND POWER**

On completing this Unit, you should be able to

- a. Define and explain work done by a force, and state units of work
- b. Identify situations in which no work is done by a force. Calculate work done under different situations
- c. Define and explain the concept of energy, and state its units.
- d. Name different forms of energy, and explain the principle of conservation of energy. Also explain the principle of conservation of work and energy.
- e. Explain kinetic and potential energies, and using the principle of conservation of energy, convert one form into the other. Solve related problems.
- f. Define and explain power, and state its units. Solve related problems.
- g. Solve problems related to the above three machines.

**Do all the problems at the end of the unit.**

### **UNIT 4: HEAT OR THERMAL ENERGY**

On completing this unit, you should be able to

- a. Describe heat as thermal energy and state its units.
- b. Describe and explain giving examples the three modes of heat transfer
- c. Describe temperature, and distinguish between heat and temperature.
- d. Explain the Celsius and the Kelvin (Absolute) scales of temperature, and convert one into the other.
- e. Define and explain (specific) heat capacity, and state its units.
- f. Discuss the implication of the large heat capacity of water and how it modifies the climate of coastal areas.
- g. State the relation between the mass of a substance, its heat capacity, change in temperature, and the amount of heat. Solve related problems.
- h. Define change of state (phase); and explain latent heat of fusion and latent heat of vaporization. State their units.
- i. Use the above definitions to calculate the amount of heat absorbed or given off when a substance undergoes a change in phase.
- j. State and explain the principle of thermal equilibrium and that of heat exchange (calorimetry), and use the principle to solve related problems.

Explain thermal expansion of substances and the coefficients of linear thermal expansion,  $\alpha$ , and volume expansion,  $\beta$ . State the relation between the two.

Solve related problems

**Do all the problems at the end of the unit.**

### **FINAL EXAM**